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Chemical Reaction
Answer

a chemical reaction in
which two or more

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Simple substances
combine to form a
new, more complex
substance

decomposition a
chemical reaction
that breaks down
compounds into
simpler products

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Chapter 6 - Chemical
Reactions ...
Lesson 2.7.1: Physical

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Science – Chemical
Reactions Part 2 H.
Turngren, Minnesota
Literacy Council,
2014 p.6 GED Science
Curriculum SCIENCE
Same Atoms, New
Bonds The reactants
and products in a
chemical reaction
contain the same
atoms, but they are
rearranged during
the reaction.

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Lesson 2.7.1: Physical Science Chemical Reactions Part 2

A change in the state of matter in which the material changes into an entirely different material with properties that are different from the original material. A change in the state of matter in which the

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material is still the same just in a different form. Tags:

Physical Science -
Chemical Reactions
Quiz - Quizizz

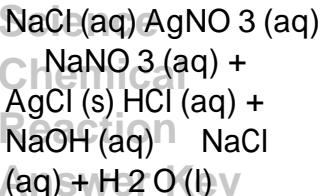
A chemical equation describes what happens in a chemical reaction. The equation identifies the reactants (starting

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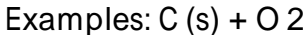
materials) and products (resulting substance), the formulas of the participants, the phases of the participants (solid, liquid, gas), and the amount of each substance.

Lesson 2.7.2: Physical
Science Chemical
Reactions Part 3 ...

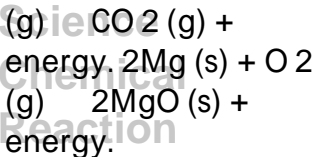
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Combustion reactions. A single element or compound combines with oxygen gas releasing energy. This rapid oxidation is called burning.



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Answer Key

Types of Chemical
Reactions - Science
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Chemical Reactions
ANSWER KEY 1.

Watch the video and
then complete the
chart. Teacher notes

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on next page! Colors:

A = Red, B = Blue, C =
Green, D = Yellow 2.

Use colored pencils
to circle the common
atoms or compounds
on each side of the
equations to help you
determine the type of
reaction it illustrates.
Use the code below
to classify each ...

Chemical Reactions

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the chapter.

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Evidence of a
chemical reaction
includes. All of the
above.(the
production of a gas, a
change in
temperature, and a

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color change.) ...

calcium fluoride. THIS
SET IS OFTEN IN
FOLDERS WITH...

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33 terms.

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Questions ...

Name _____ Chapter
7 Class _____ Date

_____ Chemical
Reactions Section 7.2
Types of Reactions
(pages 199–205)

This section discusses
how chemical
reactions are

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Science
classified into
different types.

Chemical
Reading Strategy
(page 199)

Answers Key
Previewing Skim the
section and begin a
concept map like the
one below that
identifies types of
reactions with a
general form.

Types of chemical
reactions document -

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“ KABOOM!! A chemical reaction has gone off in the science lab. Your teacher, Mrs. Beaker, needs your help. You must use your knowledge of chemical reactions to find the clues and decode them so you can escape before another chemical

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reaction occurs! Your school is counting on you. ” Immerse your students in a fun, engaging, [...]

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Chemical Reactions

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Chemical reactions are an integral part of technology, of culture, and indeed of life itself. Burning fuels, smelting iron, making glass and pottery, brewing beer, and making wine and cheese are among many examples of activities incorporating

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Chemical reactions
that have been
known and used for
thousands of years.

Chemical reactions
abound in the
geology of Earth, in
the atmosphere and
oceans ...

chemical reaction |
Definition, Equations,
Examples ...

X Your answer: For

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Science or practice,
print a copy of this
quiz at the Chemistry:
Chemical Reactions
webquest print page.

About this quiz: All
the questions on this
quiz are based on
information that can
be found at
Chemistry: Chemical
Reactions .

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Chemistry: Chemical
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9.00 g iron (III) oxide powder is reacted with 4.00 g of aluminum metal to produce molten metallic iron and aluminum oxide powder. (a) Write a balanced chemical reaction describing this process ...

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9.00 g iron (III) oxide powder is reacted with 4.00 g of ...

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The chemical

reaction $\text{Ca(OH)}_2 + \text{H}_2\text{SO}_4$

$\text{CaSO}_4 + 2$

H_2O is a: a. synthesis

reaction; b.

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decomposition
reaction; c. single
displacement
reaction; d. double
displacement
reaction; e.
combustion reaction

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